

How to solve stoichiometry problems:

1. Write a balanced reaction
2. Determine the moles of the reactant or product that you are given
 - a. If given volume use $22.4\text{L} = 1 \text{ mole}$
 - b. Or if given mass, use: molar mass = 1 mole
 - c. Or if given number of molecules, use: $1 \text{ mole} = 6.022 \times 10^{23}$ molecules
3. Use proportions given by the coefficients in the reaction to find out how many moles of the other reactant/product was used/made.
4. Now that you have moles of the substance you are asked to find, convert to the units (liters, grams, molecules) you are asked to find using the formulas in #2 above.